

Synthesis and Structures of Aluminum and Magnesium Complexes of Tetraimidophosphates and Trisamidothiophosphates: EPR and DFT Investigations of the Persistent Neutral Radicals $\{\text{Me}_2\text{Al}[(\mu\text{-NR})(\mu\text{-N}^t\text{Bu})\text{P}(\mu\text{-N}^t\text{Bu})_2]\text{Li}(\text{THF})_2\}^\bullet$ (R = SiMe₃, ^tBu)

Andrea Armstrong, Tristram Chivers,* Heikki M. Tuononen, and Masood Parvez

Department of Chemistry, University of Calgary, Calgary, Alberta, Canada T2N 1N4

Received May 3, 2005

Reactions of (RNH)₃PNSiMe₃ (**3a**, R = ^tBu; **3b**, R = Cy) with trimethylaluminum result in the formation of {Me₂Al[(μ-N^tBu)(μ-NSiMe₃)P(NH^tBu)₂]} (**4**) and the dimeric trisimidometaphosphate {Me₂Al[(μ-NCy)(μ-NSiMe₃)P(μ-NCy)₂P(μ-NCy)(μ-NSiMe₃)]AlMe₂} (**5a**), respectively. The reaction of SP(NH^tBu)₃ (**2a**) with 1 or 2 equiv of AlMe₃ yields {Me₂Al[(μ-S)(μ-N^tBu)P(NH^tBu)₂]} (**7**) and {Me₂Al[(μ-S)(μ-N^tBu)P(μ-NH^tBu)(μ-N^tBu)]AlMe₂} (**8**), respectively. Metalation of **4** with ^tBuLi produces the heterobimetallic species {Me₂Al[(μ-N^tBu)(μ-NSiMe₃)P(μ-NH^tBu)(μ-N^tBu)]Li(THF)₂} (**9a**) and {[Me₂Al][Li]₂[P(N^tBu)₃(NSiMe₃)]} (**10**) sequentially; in THF solutions, solvation of **10** yields an ion pair containing a spirocyclic tetraimidophosphate monoanion. Similarly, the reaction of (^tBuNH)₃PN^tBu with AlMe₃ followed by 2 equiv of ^tBuLi generates {Me₂Al[(μ-N^tBu)₂P(μ₂-N^tBu)₂(μ₂-THF)]Li(THF)₂} (**11a**). Stoichiometric oxidations of **10** and **11a** with iodine yield the neutral spirocyclic radicals {Me₂Al[(μ-NR)(μ-N^tBu)P(μ-N^tBu)₂]Li(THF)₂}[•] (**13a**, R = SiMe₃; **14a**, R = ^tBu), which have been characterized by electron paramagnetic resonance spectroscopy. Density functional theory calculations confirm the retention of the spirocyclic structure and indicate that the spin density in these radicals is concentrated on the nitrogen atoms of the PN₂Li ring. When **3a** or **3b** is treated with 0.5 equiv of dibutylmagnesium, the complexes {Mg[(μ-N^tBu)(μ-NH^tBu)P(NH^tBu)(NSiMe₃)₂]} (**15**) and {Mg[(μ-NCy)(μ-NSiMe₃)P(NHCy)₂]} (**16**) are obtained, respectively. The addition of 0.5 equiv of MgBu₂ to **2a** results in the formation of {Mg[(μ-S)(μ-N^tBu)P(NH^tBu)₂]} (**17**), which produces the hexameric species {[MgOH]-(μ-S)(μ-N^tBu)P(NH^tBu)₂}]₆ (**18**) upon hydrolysis. Compounds **4**, **5a**, **7–11a**, and **15–17** have been characterized by multinuclear (¹H, ¹³C, and ³¹P) NMR spectroscopy and, in the case of **5a**, **9a**·2THF, **11a**, and **18**, by X-ray crystallography.

Introduction

One of the more active areas of main-group chemistry research over the past decade has been the synthesis and characterization of imido analogues of simple oxoanions such as carbonate, silicate, and phosphate.¹ The underlying principle of this work is that the oxo units [O]²⁻ of these species are formally isoelectronic² with the imido [NR]²⁻ (R = H, alkyl, aryl) group, and thus, a new class of p-block polyanions can be generated by substituting some, or all, of the oxygen atoms by imido groups. The chemical and

physical properties of these polyimido anions are dependent, in large part, upon the ratio of imido/oxo units and the size of the imido substituent.¹

Numerous analogues of phosphorus oxoanions have been prepared,³ including the trisimido(meta)phosphate⁴ [P(NR)₃]⁻ and tetraimido(ortho)phosphate [P(NR)₃(NR')]³⁻ (R = R' =

* To whom correspondence should be addressed. Telephone: (403) 220-5741. Fax: (403) 289-9488. E-mail: chivers@ucalgary.ca.

(1) For recent reviews, see (a) Brask, J. K.; Chivers, T. *Angew. Chem., Int. Ed.* **2001**, *40*, 3960–3976. (b) Aspinall, G. M.; Copsey, M. C.; Leedham, A. P.; Russell, C. A. *Coord. Chem. Rev.* **2002**, *227*, 217–232.

(2) The original definition of “isoelectronic” (isosteric) compounds restricts the term to species having the same number of atoms and the same number of electrons (Langmuir, I. *J. Am. Chem. Soc.* **1919**, *41*, 1543). Thus, according to this definition, N³⁻ is isoelectronic with O²⁻. However, nitrides are compositionally different from oxides as a result of the different charges on the anions.

(3) Chivers, T. *Top. Curr. Chem.* **2003**, *229*, 143–159.

(4) See, for example, (a) Niecke, E.; Frost, M.; Nieger, M.; von der Gönna, V.; Ruban, A.; Schoeller, W. W. *Angew. Chem., Int. Ed. Engl.* **1994**, *33*, 2111. (b) Scherer, O. J.; Quintus, P.; Kaub, J.; Sheldrick, W. S. *Chem. Ber.* **1987**, *120*, 1463. (c) Lief, G. R.; Carrow, C. J.; Stahl, L.; Staples, R. J. *Organometallics* **2001**, *20*, 1629.

naphthyl;⁵ R = ^tBu, R' = SiMe₃⁶) trianions. Heteroleptic systems, such as the trisimidophosphate and -thiophosphate trianions [EP(NR)₃]³⁻ (E = O,⁷ S⁸), have also been reported recently, and the first transition-metal derivatives of the sulfur-containing ligand have been characterized.⁹

Previously, we have thoroughly investigated the reactivities of a series of trisaminophosphates, OP(NH^tBu)₃ (**1**),^{6,8} SP-(NHR)₃ (**2a**, R = ^tBu; **2b**, R = ⁱPr; **2c**, R = *p*-tol),^{8,10} and (RNH)₃PNSiMe₃ (**3a**, R = ^tBu; **3b**, R = Cy),^{6,11} toward alkylolithium reagents. It was found that the reactivity of these aminophosphate species increases across the series **1** < **2** < **3**. Specifically, 3-fold deprotonation of **1**, **2a**, or **2b** is not possible using *n*-butyllithium, whereas it does occur for **2c** and is readily effected for both **3a** and **3b**. More recently, we have prepared a series of zinc imidophosphates by the reactions of **1**, **2a**, **3a**, and **3b** with dimethylzinc.⁷ Interestingly, the reaction of **1** with ZnMe₂ generated several products via at least two distinct reaction pathways, whereas the reaction of **3a** and dimethylzinc produced only {MeZn-[(μ-N^tBu)(μ-NSiMe₃)P(NH^tBu)₂]}₂. In addition, two zinc complexes were isolated that contained the trisimidophosphate trianion [OP(N^tBu)₃]³⁻, which had not been observed in earlier reactions between **1** and *n*-butyllithium, trimethylaluminum, or lithium aluminum hydride.¹²

The only known main-group metal complexes of **2** and **3** are their lithium derivatives. In view of the current interest in aluminophosphates,¹³ the development of synthetic routes to a family of polyimidoaluminophosphates is certainly a worthwhile endeavor. To this end, we have examined the reactions of **2a** and **3** with trimethylaluminum. While it is not possible to effect the 3-fold deprotonation of **1** with trimethylaluminum,¹² the possibility of preparing trisimido-(thio)- or tetraimidoaluminophosphates remains. Reactions of **1–3** with MgBu₂ have also been investigated in order to gain a better understanding of the reactivity of these amido reagents toward metal alkyls.

Experimental Section

Reagents and General Procedures. All experiments were carried out under an argon atmosphere using standard Schlenk techniques. Toluene, *n*-hexane, diethyl ether, and tetrahydrofuran (THF) were dried over Na/benzophenone, distilled, and stored over molecular sieves prior to use. Trimethylaluminum (2.0 M solution in toluene), ⁿBuLi (2.5 M solution in hexanes), and MgBu₂ (1.0

solution in heptane) were used as received from Aldrich. Iodine was sublimed prior to use. (RNH)₃PNSiMe₃ (R = ^tBu, Cy),⁶ SP-(NH^tBu)₃,⁸ and OP(NH^tBu)₃⁸ were prepared as described previously.

Instrumentation. ¹H, ⁷Li, ¹³C, and ³¹P NMR spectra were collected on a Bruker DRX-400 spectrometer with chemical shifts reported relative to Me₄Si in CDCl₃ (¹H and ¹³C), LiCl in D₂O (⁷Li), and 85% H₃PO₄ in D₂O (³¹P). All spectra were collected at 22 °C. Infrared spectra were recorded as Nujol mulls on KBr plates using a Nicolet Nexus 470 FTIR spectrometer in the range 400–4000 cm⁻¹. Electron paramagnetic resonance (EPR) spectra were recorded at 22 °C on a Bruker EMX 113 spectrometer; spectral simulations were carried out using XEMR¹⁴ and WINEPR SimFonia.¹⁵ Elemental analyses were provided by the Analytical Services Laboratory, Department of Chemistry, University of Calgary.

Preparation of {Me₂Al[(μ-N^tBu)(μ-NSiMe₃)P(NH^tBu)₂]} (4). A solution of AlMe₃ (0.60 mL, 2.0 M, 1.2 mmol) in toluene was added to a clear, colorless solution of (tBuNH)₃PNSiMe₃ (0.404 g, 1.21 mmol) in hexane (20 mL) at 22 °C. After 3 h, the solvent was removed in vacuo, and the product was washed with hexane (5 mL), leaving **4** as an off-white powder (0.339 g, 0.868 mmol, 72%). ¹H NMR (C₆D₆, δ): 1.8 (br, 2 H, NH), 1.29 (d, 9 H, N^tBu), 1.16 (d, 18 H, NH^tBu), 0.34 (s, 9 H, SiMe₃), -0.22 (s, 6 H, AlMe₂). ¹³C NMR (C₆D₆, δ): 51.76 [d, NCM₃, ²J(¹³C-³¹P) = 10.9 Hz], 50.63 [d, NHCMe₃, ²J(¹³C-³¹P) = 15.2 Hz], 33.01 [d, NCM₃, ³J(¹³C-³¹P) = 34.2 Hz], 32.30 [d, NHCMe₃, ³J(¹³C-³¹P) = 15.5 Hz], 3.57 [d, SiMe₃, ³J(¹³C-³¹P) = 14.9 Hz], 1.37 (s, AlMe₂). ³¹P{¹H} NMR (C₆D₆, δ): 5.6 (s). IR (Nujol): 3404 cm⁻¹ (ν N-H). Anal. Calcd for C₁₇H₄₄N₄AlPSi: C, 52.27; H, 11.35; N, 14.34. Found: C, 52.16; H, 12.03; N, 13.92.

Preparation of {Me₂Al[(μ-NCy)(μ-NSiMe₃)P(μ-NCy)₂P(μ-NCy)(μ-NSiMe₃)]AlMe₂} (5a). A solution of AlMe₃ (0.35 mL, 2.0 M, 0.70 mmol) in toluene was added to a clear, colorless solution of (CyNH)₃PNSiMe₃ (0.286 g, 0.693 mmol) in hexane (20 mL) at 22 °C. The reaction mixture was stirred for 15 h; the volume of the solution was then reduced to 1 mL and stored at -18 °C. After 48 h, colorless crystals of **5a** were obtained (0.226 g, 0.306 mmol, 88.3%). ¹H NMR (C₆D₆, δ): 0.8–1.9 (m, Cy), 0.22 (s, 18 H, SiMe₃), -0.37 (s, 6 H, AlMe₂). ¹³C NMR (C₆D₆, δ): 51.68 (s, CMe₃), 50.05 (s, CMe₃), 37.78 [d, NCCH₂, ³J(¹³C-³¹P) = 35.6 Hz], 36.91 [d, NCCH₂, ³J(¹³C-³¹P) = 19.2 Hz], 36.78 [d, NCCH₂, ³J(¹³C-³¹P) = 18.4 Hz], 26.18 (s, NCCH₂CH₂), 25.94 (s, NCCH₂CH₂), 25.78 (s, *p*-CH₂), 2.67 [d, SiMe₃, ³J(¹³C-³¹P) = 14.0 Hz], -6.81 (s, AlMe₂). ³¹P{¹H} NMR (C₆D₆, δ): 17.1 (s). CHN analyses gave high C values owing to difficulties with the complete removal of CyNH₂.

Preparation of {Me₂Al[(μ-S)(μ-N^tBu)P(NH^tBu)₂]} (7). A solution of AlMe₃ (2.25 mL, 2.0 M, 4.5 mmol) in toluene was added to a clear, colorless solution of SP(NH^tBu)₃ (1.24 g, 4.44 mmol) in toluene (25 mL) at 22 °C. After 24 h, the solvent was removed in vacuo, leaving a colorless oil; after storage for 12 h, the product had solidified, giving **7** as a white powder (1.25 g, 3.73 mmol, 82.8%). ¹H NMR (C₆D₆, δ): 1.23 (s, 9 H, N^tBu), 1.18 (s, 18 H, NH^tBu), -0.24 [d, ⁴J(¹H-³¹P) = 1.8 Hz, 6 H, Me₂Al]. ¹³C{¹H} NMR (C₆D₆, δ): 54.32 [d, CMe₃, ²J(¹³C-³¹P) = 37.2 Hz], 52.44 [d, CMe₃, ²J(¹³C-³¹P) = 9.0 Hz], 33.03 [d, CMe₃, ³J(¹³C-³¹P) = 34.7 Hz], 31.42 [d, CMe₃, ³J(¹³C-³¹P) = 18.6 Hz], -5.20 (br, Me₂-

(5) Raithby, P. R.; Russell, C. A.; Steiner, A.; Wright, D. S. *Angew. Chem., Int. Ed. Engl.* **1997**, *36*, 649–650.

(6) Armstrong, A.; Chivers, T.; Krahn, M.; Parvez, M.; Schatte, G. *Chem. Commun.* **2002**, 2332–2333.

(7) Armstrong, A.; Chivers, T.; Krahn, M.; Parvez, M. *Can. J. Chem.* **2005**, *83*, in press.

(8) Chivers, T.; Krahn, M.; Schatte, G.; Parvez, M. *Inorg. Chem.* **2003**, *42*, 3994–4005.

(9) Rufanov, K. A.; Ziemer, B.; Meisel, M. *Dalton Trans.* **2004**, 3808–3809.

(10) Chivers, T.; Krahn, M.; Parvez, M.; Schatte, G. *Chem. Commun.* **2001**, 1992–1993.

(11) Armstrong, A.; Chivers, T.; Parvez, M.; Boeré, R. T. *Inorg. Chem.* **2004**, *43*, 3453–3460.

(12) Blais, P.; Chivers, T.; Krahn, M.; Schatte, G. *J. Organomet. Chem.* **2002**, *646*, 107–112.

(13) For a recent review, see Kimura, T. *Microporous Mesoporous Mater.* **2005**, *77*, 97–107.

(14) Eloranta, J. *XEMR*, version 0.7; University of Jyväskylä: Jyväskylä, Yliopisto, Finland.

(15) *WINEPR SimFonia*, version 1.25; Bruker Analytische Messtechnik GmbH: Karlsruhe, Germany, 1996.

Al). $^{31}\text{P}\{^1\text{H}\}$ NMR (C_6D_6 , δ): 35.6 (s). Anal. Calcd for $\text{C}_{14}\text{H}_{35}\text{N}_3$ -PSAl: C, 50.12; H, 10.52; N, 12.53. Found: C, 49.64; H, 10.97; N, 12.63.

Preparation of $\{\text{Me}_2\text{Al}[(\mu\text{-S})(\mu\text{-N}^t\text{Bu})\text{P}(\mu\text{-NH}^t\text{Bu})(\mu\text{-N}^t\text{Bu})\text{-AlMe}_2]\}$ (8). A solution of AlMe_3 (4.8 mL, 2.0 M, 9.6 mmol) in toluene was added to a clear, colorless solution of $\text{SP}(\text{NH}^t\text{Bu})_3$ (1.30 g, 4.65 mmol) in toluene (20 mL) at 22 °C. After 40 h, the solvent was removed in vacuo, leaving **8** as a white powder (1.74 g, 4.44 mmol, 95.5%). ^1H NMR (C_6D_6 , δ): 3.30 (br, d, NH), 1.33 [d, $^4J(^1\text{H}\text{-}^{31}\text{P}) = 0.9$ Hz, 9 H, N^tBu], 1.23 [d, $^4J(^1\text{H}\text{-}^{31}\text{P}) = 0.7$ Hz, 9 H, N^tBu], 1.20 [d, $^4J(^1\text{H}\text{-}^{31}\text{P}) = 0.9$ Hz, 9 H, NH^tBu], -0.25 (s, 3H, Me_2Al), -0.31 (s, 3 H, Me_2Al), -0.39 (s, 3H, Me_2Al), -0.54 (s, 3 H, Me_2Al). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6 , δ): 59.09 [d, CMe_3 , $^2J(^{13}\text{C}\text{-}^{31}\text{P}) = 12.9$ Hz], 54.99 [d, CMe_3 , $^2J(^{13}\text{C}\text{-}^{31}\text{P}) = 39.3$ Hz], 53.53 [d, CMe_3 , $^2J(^{13}\text{C}\text{-}^{31}\text{P}) = 8.8$ Hz], 33.23 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 34.0$ Hz], 32.33 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 31.2$ Hz], 30.05 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 17.9$ Hz], -5.20 (br, Me_2Al), -6.09 (br, Me_2Al). $^{31}\text{P}\{^1\text{H}\}$ NMR (C_6D_6 , δ): 40.9 (s). Anal. Calcd for $\text{C}_{16}\text{H}_{40}\text{N}_3$ -PSAl $_2$: C, 49.08; H, 10.30; N, 10.73. Found: C, 49.05; H, 10.72; N, 10.97.

Preparation of $\{\text{Me}_2\text{Al}[\text{Li}]\{\text{P}(\text{N}^t\text{Bu})_2(\text{NH}^t\text{Bu})(\text{NSiMe}_3)\}\}$ (9a). A solution of AlMe_3 (0.75 mL, 2.0 M, 1.50 mmol) in toluene was added to a stirred solution of $(^t\text{BuNH})_3\text{PNSiMe}_3$ (0.502 g, 1.50 mmol) in hexane (20 mL) at 22 °C. After 1.5 h, $^t\text{BuLi}$ (0.60 mL, 1.50 mmol) was added, resulting in a white slurry. After an additional 3 h, the solvent was removed in vacuo, leaving **9a** as a white powder (0.582 g, 1.47 mmol, 98%). Colorless X-ray quality crystals of $\{\text{Me}_2\text{Al}[(\mu\text{-N}^t\text{Bu})(\mu\text{-NSiMe}_3)\text{P}(\mu\text{-NH}^t\text{Bu})(\mu\text{-N}^t\text{Bu})\text{Li}(\text{THF})_2]\}$ (**9a**·2THF) were obtained from a THF/hexane mixture at -18 °C. ^1H NMR (d_8 -THF, δ): 1.35 (s, 9 H, ^tBu), 1.31 (s, 9 H, ^tBu), 1.29 (s, 9 H, ^tBu), 0.13 (s, 9 H, SiMe_3), -0.89 (d, 6 H, $\text{Me}_2\text{-Al}$). $^{13}\text{C}\{^1\text{H}\}$ NMR (d_8 -THF, δ): 52.39 (s, CMe_3), 51.14 [d, CMe_3 , $^2J(^{13}\text{C}\text{-}^{31}\text{P}) = 36.3$ Hz], 36.70 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 38.7$ Hz], 36.09 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 41.1$ Hz], 34.11 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 29.1$ Hz], 33.50 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 19.2$ Hz], 14.59 (s, SiMe_3), 4.26 (s, AlMe_2). $^{31}\text{P}\{^1\text{H}\}$ NMR (d_8 -THF, δ): 8.2 (s). Anal. Calcd for $\text{C}_{17}\text{H}_{43}\text{N}_4\text{LiAlSiP}$: C, 51.49; H, 10.93; N, 14.13. Found: C, 50.90; H, 10.73; N, 14.31.

Preparation of $\{\text{Me}_2\text{Al}[\text{Li}]\{\text{P}(\text{N}^t\text{Bu})_3(\text{NSiMe}_3)\}\}$ (10). A solution of AlMe_3 (0.60 mL, 2.0 M, 1.2 mmol) in toluene was added to a stirred pale yellow solution of $(^t\text{BuNH})_3\text{PNSiMe}_3$ (0.404 g, 1.21 mmol) in hexane, resulting in a cloudy off-white solution. After 3h, $^t\text{BuLi}$ (0.90 mL, 2.5 M, 2.3 mmol) in hexanes was added to the reaction mixture, resulting in a white slurry. After an additional 3 h, the solvent was removed in vacuo, leaving **10** as a white powder (0.427 g, 1.06 mmol, 93%). ^1H NMR (d_8 -THF, δ): 1.32 (s, 9 H, N^tBu), 1.30 (s, 18 H, N^tBu), 0.08 (s, 9 H, SiMe_3), -0.94 (s, 6 H, AlMe_2). ^7Li (d_8 -THF, δ): 0.70 (s). ^7Li (C_6D_6 , δ): 1.44 (s), -2.71 (s). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6 , δ): 50.72 [d, NCMe_3 , $^2J(^{13}\text{C}\text{-}^{31}\text{P}) = 6.8$ Hz], 36.58 [d, NCMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 36.2$ Hz], 33.77 [d, NHCM_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 27.2$ Hz], 4.19 (s, SiMe_3), -4.61 (s, AlMe_2). $^{31}\text{P}\{^1\text{H}\}$ NMR (d_8 -THF, δ): 10.0 (s). Satisfactory CHN analyses could not be obtained because of the extremely hygroscopic nature of this compound.

Preparation of $\{\text{Me}_2\text{Al}[(\mu\text{-N}^t\text{Bu})_2\text{P}(\mu_2\text{-N}^t\text{Bu})_2(\mu_2\text{-THF})\text{Li}(\text{THF})_2]\}$ (11a). A solution of AlMe_3 (0.55 mL, 2.0 M, 1.10 mmol) in toluene was added to a stirred clear, colorless solution of $(^t\text{BuNH})_3\text{PN}^t\text{Bu}$ (0.353 g, 1.11 mmol) in hexane (15 mL), resulting in an opaque white solution. After 3 h, $^t\text{BuLi}$ (1.4 mL, 1.6 M, 2.24 mmol) in hexanes was added to the reaction mixture, resulting in a white slurry. After an additional 2.5 h, the white solid was allowed to settle, and the supernatant liquid was decanted. The product was washed with an additional 5 mL of hexane prior to the removal of

the solvent in vacuo. The product $\{\text{Me}_2\text{Al}[\text{Li}]\{\text{P}(\text{N}^t\text{Bu})_4\}\}$ was isolated as a white powder (0.339 g, 0.887 mmol, 80%). Recrystallization from THF/hexane produced colorless X-ray quality crystals of $\{\text{Me}_2\text{Al}[(\mu\text{-N}^t\text{Bu})_2\text{P}(\mu_2\text{-N}^t\text{Bu})_2(\mu_2\text{-THF})\text{Li}(\text{THF})_2]\}$ (**11a**). ^1H NMR (d_8 -THF, δ): 1.34 (s, 36 H, N^tBu), -0.96 (s, 6 H, AlMe_2). ^7Li (d_8 -THF, δ): 1.23 (s). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6 , δ): 50.90 (s, NCMe_3), 50.78 [d, NCMe_3 , $^2J(^{13}\text{C}\text{-}^{31}\text{P}) = 8.3$ Hz], 36.41 [d, NCMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 34.9$ Hz], 34.45 [d, NHCM_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 25.4$ Hz], -4.0 (br, s, AlMe_2). $^{31}\text{P}\{^1\text{H}\}$ NMR (d_8 -THF, δ): 12.4 (s). Satisfactory CHN analyses could not be obtained because of the extremely hygroscopic nature of this compound.

Preparation of $\{\text{Mg}[(\mu\text{-N}^t\text{Bu})(\mu\text{-NH}^t\text{Bu})\text{P}(\text{NH}^t\text{Bu})(\text{NSiMe}_3)_2]\}$ (15). A solution of MgBu_2 (1.56 mL, 1.0 M, 1.56 mmol) in heptane was added to a stirred solution of $(^t\text{BuNH})_3\text{PNSiMe}_3$ (1.05 g, 3.14 mmol) in hexane (15 mL) at 22 °C. After 18 h, the solvent was removed in vacuo, leaving **15** as a white solid (0.986 g, 1.42 mmol, 91%). ^1H NMR (C_6D_6 , δ): 1.9 (br, d, NH), 1.8 (br, d, NH), 1.49 (s, 9 H, N^tBu), 1.39 (s, 9 H, NH^tBu), 1.32 (s, 9 H, NH^tBu), 0.50 (s, 9 H, SiMe_3). $^{13}\text{C}\{^1\text{H}\}$ NMR (d_8 -THF, δ): 51.52 (s, CMe_3), 51.46 (s, CMe_3), 51.40 (s, CMe_3), 50.89 [d, CMe_3 , $^2J(^{13}\text{C}\text{-}^{31}\text{P}) = 3.0$ Hz], 50.63 [d, CMe_3 , $^2J(^{13}\text{C}\text{-}^{31}\text{P}) = 15.0$ Hz], 35.2 (s, CMe_3), 33.4 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 30.0$ Hz], 33.2 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 30.0$ Hz], 32.3 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 30.0$ Hz], 5.7 [d, SiMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 30.0$ Hz], 4.3 [d, SiMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 30.0$ Hz]. $^{31}\text{P}\{^1\text{H}\}$ NMR (C_6D_6 , δ): 5.2 (s). Anal. Calcd for $\text{C}_{30}\text{H}_{76}\text{N}_8\text{MgP}_2\text{Si}_2$: C, 52.11; H, 11.08; N, 16.21. Found: C, 52.20; H, 10.19; N, 15.51.

Preparation of $\{\text{Mg}[(\mu\text{-NCy})(\mu\text{-NSiMe}_3)\text{P}(\text{NHCy})_2]\}$ (16). A solution of MgBu_2 (0.82 mL, 0.82 mmol) in heptane was added to a stirred solution of $(\text{CyNH})_3\text{PNSiMe}_3$ (0.677 g, 1.64 mmol) in hexane (20 mL) at 22 °C. After 20 h, the solvent was removed in vacuo, leaving **16** as an off-white solid (0.540 g, 0.637 mmol, 78%). ^1H NMR (C_6D_6 , δ): 1.0–3.2 (br, m, Cy), 0.40 (s, 9 H, SiMe_3). $^{13}\text{C}\{^1\text{H}\}$ NMR (d_8 -THF, δ): 52.71 (s, CN), 50.03 [d, CNH, $^2J(^{13}\text{C}\text{-}^{31}\text{P}) = 30.4$ Hz], 40.26 (s, CH_2CN), 37.70 [d, CH_2CNH , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 80.0$ Hz], 26.73 (s, CHCH_2), 26.31 (br, s, CH_2CHCH_2), 4.34 (s, SiMe_3). $^{31}\text{P}\{^1\text{H}\}$ NMR (C_6D_6 , δ): 19.6 (s). Anal. Calcd for $\text{C}_{42}\text{H}_{88}\text{N}_8\text{P}_2\text{Si}_2\text{Mg}$: C, 59.51; H, 10.46; N, 13.22. Found: C, 59.46; H, 10.88; N, 12.53.

Preparation of $\{\text{Mg}[(\mu\text{-S})(\mu\text{-N}^t\text{Bu})\text{P}(\text{NH}^t\text{Bu})_2]\}$ (17). A solution of MgBu_2 (2.0 mL, 1.0 M, 2.0 mmol) was added to a stirred solution of $\text{SP}(\text{NH}^t\text{Bu})_3$ (1.125 g, 4.026 mmol) in toluene (20 mL) at 22 °C, resulting in a clear, colorless solution. After 20 h, the solvent was removed in vacuo, leaving a colorless oil. After storage for 24 h, **17** was obtained as a white powder (1.077 g, 1.850 mmol, 92.5%). ^1H NMR (C_6D_6 , δ): 2.03 (br, d, NH), 1.42 (s, 9H, N^tBu), 1.30 (s, 18H, NH^tBu). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6 , δ): 53.32 [d, CMe_3 , $^2J(^{13}\text{C}\text{-}^{31}\text{P}) = 46.7$ Hz], 51.76 [d, CMe_3 , $^2J(^{13}\text{C}\text{-}^{31}\text{P}) = 8.8$ Hz], 34.73 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 41.0$ Hz], 31.63 [d, CMe_3 , $^3J(^{13}\text{C}\text{-}^{31}\text{P}) = 16.9$ Hz]. $^{31}\text{P}\{^1\text{H}\}$ NMR (C_6D_6 , δ): 38.8 (s). Anal. Calcd for $\text{C}_{24}\text{H}_{58}\text{N}_6\text{P}_2\text{S}_2\text{Mg}$: C, 49.60; H, 10.06; N, 14.46. Found: C, 50.89; H, 10.62; N, 13.69.

Storage of a solution of **17** in diethyl ether at -18 °C for one week yielded colorless crystals of $\{\text{MgOH}\}[(\mu\text{-S})(\mu\text{-N}^t\text{Bu})\text{P}(\text{NH}^t\text{Bu})_2]\}_6$ (**18**).

X-ray Analyses. Colorless crystals of **5a**, **9a**·2THF, **11a**, and **18** were coated with Paratone 8277 oil and mounted on a glass fiber. All measurements were made on a Nonius Kappa CCD diffractometer using graphite-monochromated Mo K α radiation. Crystallographic data are summarized in Table 1. The structures were solved by direct methods¹⁶ and refined by full-matrix least-

(16) Altomare, A.; Cascarano, M.; Giacovazzo, C.; Guagliardi, A. *J. Appl. Crystallogr.* **1993**, *26*, 343.

Table 1. Crystallographic Data

	5a	9a·2THF	11a	18
formula	C ₃₄ H ₇₄ Al ₂ N ₆ P ₂ Si ₂	C ₂₅ H ₅₉ AlLiN ₄ O ₂ PSi	C ₃₄ H ₇₄ AlLi ₂ N ₄ O ₄ P	C ₄₀ H ₁₀₀ Mg ₃ N ₉ O ₄ P ₃ S ₃
fw, g	739.07	540.74	674.80	1033.31
space group	P2 ₁ /c	P2 ₁ /c	P2 ₁ /c	P1̄
a, Å	20.400(6)	10.209(3)	10.876(2)	14.993(4)
b, Å	16.813(4)	17.616(4)	12.901(3)	15.285(4)
c, Å	12.865(11)	18.975(6)	29.422(7)	15.535(4)
α, deg	90	90	90	96.188(14)
β, deg	92.31(3)	91.218(9)	91.541(13)	102.701(16)
γ, deg	90	90	90	117.474(14)
V, Å ³	4409(4)	3411.7(17)	4126.7(16)	2989.9(14)
Z	4	4	4	2
T, K	173(2)	173(2)	100(2)	173(2)
λ, Å	0.710 73	0.710 73	0.710 73	0.710 73
d _{calcd.} , g cm ⁻³	1.113	1.053	1.086	1.148
μ, mm ⁻¹	0.223	0.17	0.125	0.278
F(000)	1616	1192	1488	1128
R ^a	0.044,	0.081	0.23	0.042
R _w ^b	0.107	0.203	0.55	0.096

^a $R = [\sum ||F_o| - |F_c||] / [\sum |F_o|]$ for reflections with $I > 2.00\sigma(I)$. ^b $R_w = \{[\sum w(F_o^2 - F_c^2)^2] / [\sum w(F_o^2)^2]\}^{1/2}$ for all reflections.

squares methods using SHELXL-97.¹⁷ Hydrogen atoms were included at geometrically idealized positions and were not refined; the non-hydrogen atoms were refined anisotropically. A cyclohexyl ring in **5a** was disordered in one of the molecules; this was included in the refinements using partial site occupancy factors. Only some of the non-hydrogen atoms in **9a**·2THF were refined anisotropically. The α carbon of the N^tBu group and the silicon atom of the NSiMe₃ group chelated to the Al atom showed positional disorder with the two atoms occupying the same sites; the model was refined with the C and Si atoms assigned partial occupancies that summed to unity. Carbon atoms of both THF ligands were disordered over two sites with equal site occupancy factors. The hydrogen atom attached to N(4) was located from a difference map.

Computational Details. The structures of the model radicals {Me₂Al[(μ-NMe)(μ-NSiH₃)P(μ-NMe)₂Li(OMe)₂]}* (**13b**) and {Me₂-Al[(μ-NMe)₂P(μ-NMe)₂Li(OMe)₂]}* (**14b**) were optimized in their doublet ground states using density functional theory (DFT) and C_s and C_{2v} symmetry, respectively. A hybrid PBE0 exchange-correlation functional^{18,19} and Ahlrichs' triple-ζ valence basis set augmented by one set of polarization functions (TZVP)²⁰ were used in the optimizations. Hyperfine coupling constants (HFCCs) were then calculated by single-point calculations employing the optimized geometries, PBE0 functional, and unrestricted Kohn–Sham formalism. For P, N, Li, and Al, the single-point calculations utilized the IGLO-III basis set²¹ in a completely uncontracted form and were augmented with one additional steep s function as well as f-polarization functions from the cc-pVTZ basis set.²¹ For Si, C, and H, the calculations utilized the cc-pVTZ basis set²¹ in its standard form. A pruned (99 590) integration grid was used in all single-point calculations to ensure numerical convergence of calculated HFCCs. All geometry optimizations were done with the Turbomole 5.7²² program package; Gaussian 03²³ was used in all single-point calculations.

(17) Sheldrick, G. M. *SHELXL-97*; University of Göttingen: Göttingen, Germany, 1997.

(18) (a) Perdew, J. P.; Burke, K.; Ernzerhof, M. *Phys. Rev. Lett.* **1996**, *77*, 3865–3868. (b) Perdew, J. P.; Burke, K.; Ernzerhof, M. *Phys. Rev. Lett.* **1997**, *78*, 1396–1396 (E).

(19) Perdew, J. P.; Ernzerhof, M.; Burke, K. *J. Chem. Phys.* **1996**, *105*, 9982–9985.

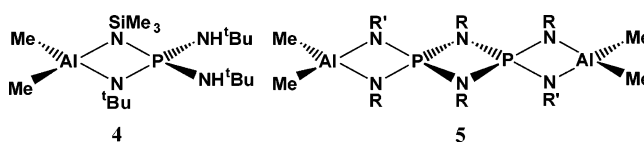
(20) The TZVP basis set was used as it is referenced in the Turbomole 5.7 internal basis set library.

(21) The basis set was taken from the EMSL basis set library: <http://www.emsl.pnl.gov/forms/basisform.html> (accessed December 2004).

(22) *TURBOMOLE*, version 5.7. R. Theoretical Chemistry Group, University of Karlsruhe: Karlsruhe, Germany, 2004.

Results and Discussion

Synthesis and Characterization of {Me₂Al[(μ-N^tBu)(μ-NSiMe₃)P(NH^tBu)₂]} (4) and {Me₂Al[(μ-NCy)(μ-NSiMe₃)P(μ-NCy)₂P(μ-NCy)(μ-NSiMe₃)]AlMe₂} (5a). The reaction of (tBuNH)₃PNSiMe₃ (**3a**) with trimethylaluminum in a 1:1 molar ratio results in the elimination of 1 equiv of methane and the formation of {Me₂Al[(μ-N^tBu)(μ-NSiMe₃)P(NH^tBu)₂]} (**4**) in good yield. On the basis of NMR data, complex **4** is assigned a monocyclic structure in which the aluminum center is chelated by the two imido groups (N^tBu and NSiMe₃) of the monoanionic ligand (formally isoelectronic with [H₂PO₄]²⁻), whereas the two NH^tBu groups are in exocyclic positions. The ¹H NMR spectrum of **4** in C₆D₆ contains two ^tBu resonances at 1.16 and 1.29 ppm (relative intensities 2:1), attributed to the NH^tBu and N^tBu groups, respectively. Resonances due to the SiMe₃ and AlMe₂ units with the appropriate intensities are observed at 0.34 and -0.22 ppm, respectively. The ¹³C NMR spectrum of **4** exhibits resonances for two types of ^tBu groups, one SiMe₃ group, and an AlMe₂ unit, consistent with this structural proposal; the ³¹P spectrum reveals a singlet at 5.6 ppm. Single crystals of **4** could not be obtained owing to its high solubility in organic solvents.



Interestingly, the reaction of the cyclohexyl derivative (CyNH)₃PNSiMe₃ (**3b**) with AlMe₃ proceeds in a different manner. In this case, elimination of 1 equiv of cyclohexylamine results in the nearly quantitative formation of {Me₂Al[(μ-NCy)(μ-NSiMe₃)P(μ-NCy)₂P(μ-NCy)(μ-NSiMe₃)]AlMe₂} (**5a**). An X-ray structure of **5a** revealed a centrosymmetric dimer in which the central feature is an essentially planar P₂N₂ ring with the two phosphorus atoms linked by NCy groups (Figure 1). Complex **5a** is the dimethylaluminum derivative of the trisimidometaphosphate monoanion [P(NCy)₂(NSiMe₃)]⁻. Two closely related com-

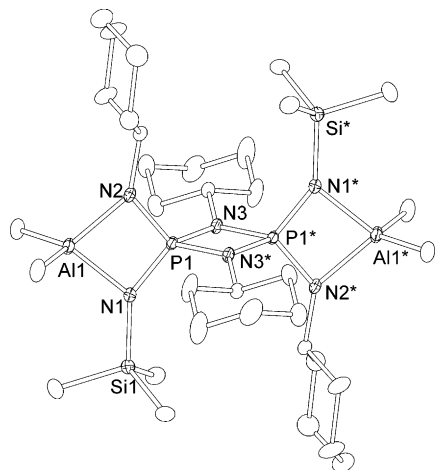


Figure 1. Thermal ellipsoid diagram of **5a** (30% probability ellipsoids). Hydrogen atoms omitted for clarity.

Table 2. Selected Bond Lengths (Å) and Angles (deg) for **5a**^a

P(1)–N(1)	1.5903(15)	P(1)–N(2)	1.5952(15)
P(1)–N(3)	1.6816(16)	Al(1)–C(4)	1.959(2)
Al(1)–C(5)	1.963(2)	Al(1)–N(1)	1.9597(16)
Al(1)–N(2)	1.9251(16)	P(1)–N(3)–P(1)*	96.55(9)
N(3)–P(1)–N(3)*	83.45(9)	N(1)–Al(1)–N(2)	76.40(6)
P(1)–N(1)–Al(1)	92.27(7)	P(1)–N(2)–Al(1)	93.41(7)

^a Symmetry transformations used to generate equivalent atoms: #1 $-x, -y, -z$; #2 $-x + 1, -y, -z + 1$.

plexes have been isolated by Stahl et al.^{4c} from the reaction of the bis(amino)cyclodiphosph(V)azanes *cis*-{(tBuNH)(RN=PN^tBu)₂(NH^tBu)} with trimethylaluminum. The general structure of these tricyclic species is shown above (**5a**, R = Cy, R' = SiMe₃; **5b**, R = ^tBu, R' = Ph; **5c**, R = ^tBu, R' = *p*-tol).

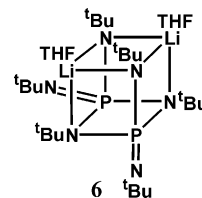
Selected bond lengths and bond angles of **5a** are shown in Table 2. Each phosphorus atom is also part of a planar four-membered PN₂Al ring that is approximately perpendicular to the P₂N₂ ring. The dimethylaluminum cations are chelated by an NSiMe₃ group and the NCy group that is not involved in the central P₂N₂ ring. The P–N distances involving these two imido ligands are very similar [P(1)–N(1) = 1.5903(15) Å; P(1)–N(2) = 1.5952(15) Å] despite their different substituents. These bond lengths are significantly shorter than those in the P₂N₂ ring [P(1)–N(3) = 1.6816(16) Å], indicating a higher formal bond order in the former. The two Al–N bond lengths also differ slightly

[Al(1)–N(1) = 1.9597(16) Å; Al(1)–N(2) = 1.9251(16) Å], with the shorter bond observed in the case of the more basic NCy ligand.

A comparison of structural parameters of **5a** and **5c**^{4c} shows the expected similarities, with P–N single bond distances of 1.701(2) Å observed within the P₂N₂ ring of **5c** [cf. 1.6816(16) Å in **5a**] and the Al–N distances in **5a** [*av* = 1.942(2) Å] slightly shorter than the average bond length in **5c** [1.958(2) Å]. Conversely, longer Al–C distances are observed in **5a** [1.961(2) Å] compared to those in **5c** [1.944(2) Å]. These minor differences can be attributed to the more electron-donating character of the cyclohexylimido substituents of **5a** as compared to the NR (R = ^tBu, *p*-tol) groups of **5c**.

The NMR spectra of **5a** in C₆D₆ are consistent with the retention of its dimeric structure in solution. In particular, the ¹³C NMR spectrum contained two sets of cyclohexyl resonances as well as signals attributed to the SiMe₃ (2.67 ppm) and AlMe₂ (–6.81 ppm) fragments. The ³¹P NMR signal for **5a** appears at 17.1 ppm. While the yields involved in the two-step syntheses of the imidometaphosphate dimers **5b** and **5c** range from moderate to excellent,^{4c,24} the preparation of **5a** requires only one step and does not involve the use of potentially explosive organic azides.

Synthesis and NMR Characterization of {Me₂Al[(μ-S)(μ-N^tBu)P(NH^tBu)₂]} (7) and {Me₂Al[(μ-S)(μ-N^tBu)P(μ-NH^tBu)(μ-N^tBu)]AlMe₂} (8). In earlier attempts⁸ to prepare a trisimido(thio)phosphate trianion [SP(NR)₃]^{3–}, it was found that the reaction of SP(NH^tBu)₃ (**2a**) with 2 equiv of *n*-butyllithium results in the extrusion of sulfur and the formation of the cubic tris(imido)metaphosphate dimer {[μ₃-Li(THF)][P(μ₃-N^tBu)₂(N^tBu)]₂} (**6**) in 15% yield. In light of the results described earlier, investigations into the reactions of trisamino(thio)phosphates SP(NHR)₃ (**2**) with AlMe₃ are of considerable interest, as three possible reaction pathways can be envisioned: (i) monodeprotonation of **2** to yield a [Me₂Al]⁺ complex of the [SP(NR)(NHR)₂][–] anion, (ii) deprotonation and loss of RNH₂ to yield a bis(imido)(thio)metaphosphate [SP(NR)₂][–] monoanion, or (iii) deprotonation and sulfur extrusion to generate a tris(imido)metaphosphate [P(NR)₃][–] complex.

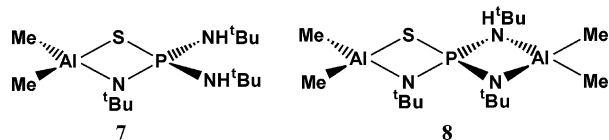


The reaction of SP(NH^tBu)₃ (**2a**) with 1 equiv of AlMe₃ results in the formation of **7**, a white powder, which has been characterized by multinuclear NMR spectroscopy. The ³¹P NMR spectrum of the product in C₆D₆ contained a single resonance at 35.6 ppm (cf. 46.3 ppm for **2a**), a strong indication that a loss of sulfur has not occurred, since this would be accompanied by a significant upfield shift as is

- (23) Frisch, M. J.; Trucks, G. W.; Schlegel, H. B.; Scuseria, G. E.; Robb, M. A.; Cheeseman, J. R.; Montgomery, Jr., J. A.; Vreven, T.; Kudin, K. N.; Burant, J. C.; Millam, J. M.; Iyengar, S. S.; Tomasi, J.; Barone, V.; Mennucci, B.; Cossi, M.; Scalmani, G.; Rega, N.; Petersson, G. A.; Nakatsuji, H.; Hada, M.; Ehara, M.; Toyota, K.; Fukuda, R.; Hasegawa, J.; Ishida, M.; Nakajima, T.; Honda, Y.; Kitao, O.; Nakai, H.; Klene, M.; Li, X.; Knox, J. E.; Hratchian, H. P.; Cross, J. B.; Bakken, V.; Adamo, C.; Jaramillo, J.; Gomperts, R.; Stratmann, R. E.; Yazyev, O.; Austin, A. J.; Cammi, R.; Pomelli, C.; Ochterski, J. W.; Ayala, P. Y.; Morokuma, K.; Voth, G. A.; Salvador, P.; Dannenberg, J. J.; Zakrzewski, V. G.; Dapprich, S.; Daniels, A. D.; Strain, M. C.; Farkas, O.; Malick, D. K.; Rabuck, A. D.; Raghavachari, K.; Foresman, J. B.; Ortiz, J. V.; Cui, Q.; Baboul, A. G.; Clifford, S.; Cioslowski, J.; Stefanov, B. B.; Liu, G.; Liashenko, A.; Piskorz, P.; Komaromi, I.; Martin, R. L.; Fox, D. J.; Keith, T.; Al-Laham, M. A.; Peng, C. Y.; Nanayakkara, A.; Challacombe, M.; Gill, P. M. W.; Johnson, B.; Chen, W.; Wong, M. W.; Gonzalez, C.; Pople, J. A. *Gaussian 03*, revision C.02; Gaussian, Inc.: Wallingford, CT, 2004.

- (24) Moser, D. F.; Carrow, C. J.; Stahl, L.; Staples, R. J. *J. Chem. Soc., Dalton Trans.* **2001**, 8, 1246–1252.

observed for **6** (^{31}P NMR: -21 ppm).⁸ The ^1H NMR spectrum displayed two *tert*-butyl signals at 1.23 and 1.18 ppm in a 1:2 ratio, along with a peak at -0.24 ppm attributed to a Me_2Al^+ cation; the corresponding signals were observed in the ^{13}C NMR spectrum. These NMR data are consistent with the formation of $\{\text{Me}_2\text{Al}[(\mu\text{-S})(\mu\text{-N}^t\text{Bu})\text{P}(\text{NH}^t\text{Bu})_2]\}$ (**7**), a monocyclic species that is isostructural with **4** and, like **4**, is the product of a simple deprotonation reaction.



The reaction of **2a** with 2 equiv of AlMe_3 yields a white powder that has been identified as the spirocyclic complex $\{\text{Me}_2\text{Al}[(\mu\text{-S})(\mu\text{-N}^t\text{Bu})\text{P}(\mu\text{-NH}^t\text{Bu})(\mu\text{-N}^t\text{Bu})]\text{AlMe}_2\}$ (**8**) on the basis of multinuclear NMR data. A single ^{31}P resonance was observed at 40.9 ppm, indicating that the P–S bond remains intact (*vide supra*). The ^1H NMR spectrum in C_6D_6 contained three *tert*-butyl resonances of equal intensities, as well as four signals between -0.54 and -0.25 ppm attributed to the methyl groups of the two inequivalent dimethylaluminum cations. The corresponding ^tBu signals were observed in the ^{13}C NMR spectrum, along with two broad resonances assigned to the Me_2Al^+ cations. The fact that the four methyl groups resonate at different ^1H NMR frequencies is attributed to the lack of planarity in both the AlSNP and AlN_2P rings, resulting in a molecule with only C_1 symmetry.

Further investigations of the reaction of **2a** with AlMe_3 as a function of stoichiometry, reaction time, and reaction temperature did not yield any products other than **7** and **8**. Similarly, attempts to form a methylaluminum complex of the amidotris(imido)phosphate dianion $[\text{P}(\text{NH}^t\text{Bu})(\text{N}^t\text{Bu})_2(\text{NSiMe}_3)]^{2-}$ by increasing the reaction time and temperature or changing the stoichiometry of the reagents were not successful; in all cases, only **4** was obtained. The nearly quantitative formation of the simple dimethylaluminum derivatives **4**, **7**, and **8** is in sharp contrast to our observations for the reactions of $\text{OP}(\text{NH}^t\text{Bu})_3$ (**1**) with AlMe_3 in which products involving monoanionic and dianionic ligands, as well as the neutral ligand **1**, were obtained.¹²

Synthesis and NMR Characterization of $\{\text{Me}_2\text{Al}[(\mu\text{-N}^t\text{Bu})(\mu\text{-NSiMe}_3)\text{P}(\mu\text{-NH}^t\text{Bu})(\mu\text{-N}^t\text{Bu})]\text{Li}(\text{THF})_2\}$ (9a**·2THF), $\{\text{Me}_2\text{Al}[\text{Li}]_2[\text{P}(\text{N}^t\text{Bu})_3(\text{NSiMe}_3)]\}$ (**10**), and $\{\text{Me}_2\text{Al}[(\mu\text{-N}^t\text{Bu})_2\text{P}(\mu\text{-N}^t\text{Bu})_2(\mu\text{-THF})][\text{Li}(\text{THF})_2]\}$ (**11a**); X-ray Structures of **9a**·2THF and **11a**.** In view of the limited reactivity of AlMe_3 toward $(^t\text{BuNH})_3\text{PNSiMe}_3$ (**3a**), deprotonation of the dimethylaluminum complex **4** with $^t\text{BuLi}$ was investigated. The treatment of **4** with 1 equiv of $^t\text{BuLi}$ resulted in the isolation of the heterobimetallic complex $\{-[\text{Me}_2\text{Al}][\text{Li}][\text{P}(\text{N}^t\text{Bu})_2(\text{NH}^t\text{Bu})(\text{NSiMe}_3)]\}$ (**9a**) in essentially quantitative yield. Recrystallization of **9a** from THF/hexane produced colorless crystals of **9a**·2THF; an X-ray structure (Figure 2) revealed that this complex exists as the spirocycle $\{\text{Me}_2\text{Al}[(\mu\text{-N}^t\text{Bu})(\mu\text{-NSiMe}_3)\text{P}(\mu\text{-NH}^t\text{Bu})(\mu\text{-N}^t\text{Bu})]\text{Li}(\text{THF})_2\}$ in which the dimethylaluminum cation is *N,N'*-chelated by the NSiMe_3 group and one N^tBu unit, while the

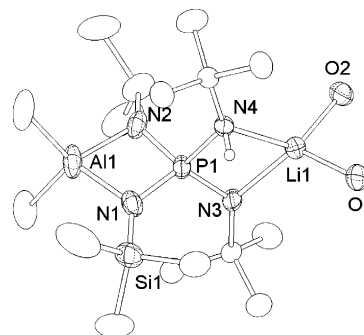


Figure 2. Thermal ellipsoid diagram of **9a**·2THF (30% probability ellipsoids). Only α -carbon atoms of the ^tBu and SiMe_3 groups are shown; only oxygen atoms of the THF molecules are shown.

lithium cation is coordinated to the remaining N^tBu and NH^tBu groups and to two molecules of tetrahydrofuran. Pertinent bond lengths and angles are given in Table 3 and are in good agreement with metrical parameters calculated for the model system $\{\text{Me}_2\text{Al}[(\mu\text{-NMe})(\mu\text{-NSiH}_3)\text{P}(\mu\text{-NHMe})(\mu\text{-NMe})]\text{Li}(\text{OME}_2)_2\}$ (**9b**·2 OME_2), particularly when the difference in the identities of the imido groups is taken into account.

The P–N distances in **9a**·2THF range from 1.588(3) to 1.716(3) Å, with identical bond lengths of 1.640(4) Å observed for the two nitrogen atoms that chelate the aluminum center. A longer P–N distance is present between the phosphorus and the *tert*-butylamido nitrogen [1.716(3) Å], while a significantly shorter distance is found for the $\text{P}=\text{N}^t\text{Bu}$ bond [1.588(3) Å], indicating the presence of localized single and double phosphorus–nitrogen bonds for N(4) and N(3), respectively. The mean Al–N distance in **9a**·2THF [1.899(4) Å] is slightly shorter than that observed in **5a** [1.9422(2) Å], presumably because of the larger formal charge on each of the chelating imido groups of **9a**·2THF as compared to those of **5a**. As expected, the Li–N bond to the monanion NH^tBu ligand [2.178(7) Å] is substantially longer than the Li–N linkage involving the dianionic N^tBu ligand [1.969(7) Å]. The angles about these two nitrogen atoms are also significantly different [$\angle\text{PN}(3)\text{Li} = 100.6(3)^\circ$; $\angle\text{PN}(4)\text{Li} = 88.9(2)^\circ$] because of their different hybridizations. In consequence, while the four-membered AlN_2P ring is essentially planar [$\angle\text{N}(1)\text{Al}(1)\text{N}(2)\text{P}(1) = -1.15(15)^\circ$], the LiN_2P ring is puckered [$\angle\text{N}(4)\text{P}(1)\text{N}(3)\text{Li}(1) = 17.9(3)^\circ$]. This X-ray structure confirmed the coordination mode of the ligand to the AlMe_2 unit of **4** that was deduced from the NMR data (*vide supra*).

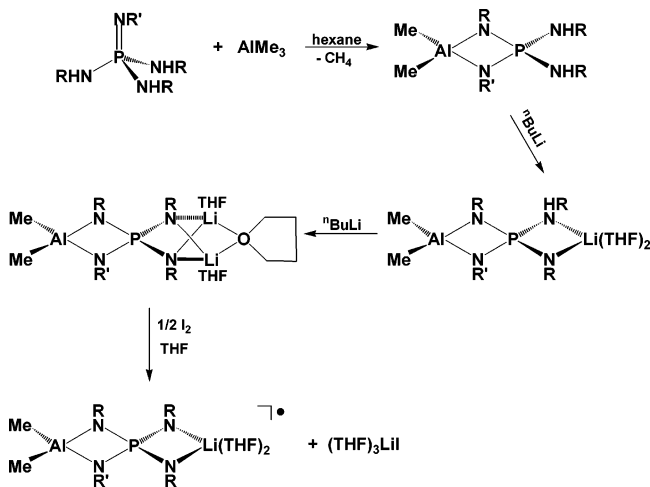
The ^1H NMR spectrum of **9a** in d_8 -THF contains three ^tBu signals of equal intensity at 1.35, 1.31, and 1.29 ppm, as well as a single SiMe_3 resonance at 0.13 ppm and a singlet at -0.89 ppm due to the two equivalent methyl groups of the AlMe_2 unit. The corresponding signals were observed in the ^{13}C NMR spectrum, while a single resonance at 8.2 ppm was observed in the ^{31}P NMR spectrum. These spectroscopic data are consistent with the retention of the spirocyclic structure of **9a**·2THF in THF solution.

In a similar manner, the dilithiated complex $\{[\text{Me}_2\text{Al}][\text{Li}]_2[\text{P}(\text{N}^t\text{Bu})_3(\text{NSiMe}_3)]\}$ (**10**) is recovered in excellent yield from the reaction of **4**, prepared *in situ*, with 2 equiv of $^t\text{BuLi}$ (Scheme 1). An analogous synthesis of the sym-

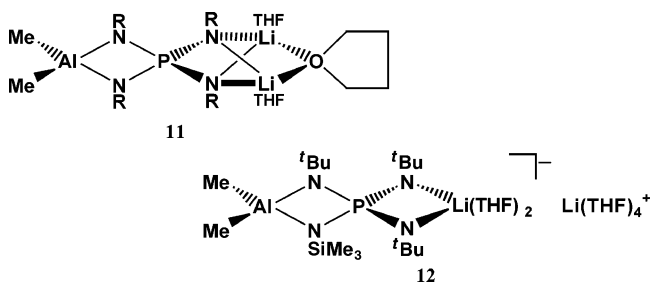
Table 3. Selected Bond Lengths (Å) and Angles (deg) for **9**

atoms	X-ray data for 9a ·2THF	calcd for 9b ·2OMe ₂	atoms	X-ray data for 9a ·2THF	calcd for 9b ·2OMe ₂
P(1)–N(1)	1.640(4)	1.642	P(1)–N(2)	1.640(4)	1.643
P(1)–N(3)	1.588(3)	1.611	P(1)–N(4)	1.716(3)	1.755
Al(1)–N(1)	1.911(5)	1.966	Al(1)–N(2)	1.886(4)	1.922
Li(1)–N(3)	1.969(7)	1.933	Li(1)–N(4)	2.178(7)	2.137
Al(1)–C(1)	1.980(5)	1.983	Al(1)–C(2)	1.989(5)	1.986
Li(1)–O(1)	1.993(8)	1.974	N(1)–Al(1)–N(2)	77.85(16)	76.8
N(3)–Li(1)–N(4)	71.6(2)	75.8	P(1)–N(4)–Li(1)	88.9(2)	86.0
P(1)–N(3)–Li(1)	100.6(3)	97.0	P(1)–N(1)–Al(1)	93.9(2)	93.4
P(1)–N(2)–Al(1)	94.9(2)	95.3	N(1)–P(1)–N(2)	93.3(2)	94.5
N(1)–P(1)–N(3)	121.63(19)	123.2	N(3)–P(1)–N(4)	94.71(16)	96.1

Scheme 1. Stepwise Deprotonation of Tris(amino)phosphates to Mixed Me₂Al⁺/Li⁺ Tetraimidophosphate Complexes and Subsequent Oxidation to Yield Persistent Spirocyclic Radicals (R = ^tBu; R' = SiMe₃, ^tBu)



metrical tetraimidophosphate trianion [P(N^tBu)₄]³⁻ as the heterobimetallic complex { [Me₂Al][Li]₂[P(N^tBu)₄] } was carried out by treating (^tBuNH)₃PN^tBu first with AlMe₃, then with ⁿBuLi. The structure of the THF adduct { Me₂Al[(μ-N^t-Bu)₂P(μ₂-N^tBu)₂(μ₂-THF)][Li(THF)₂] } (**11a**) was determined by X-ray crystallography. This complex crystallizes with a molecule of THF in its lattice, rendering the crystals unstable toward solvent loss once removed from solution. Consequently, the X-ray data are not of sufficient quality to allow for a discussion of the bond lengths and angles of this compound. However, the structural parameters are consistent with the corresponding parameters that were calculated for the model system { Me₂Al[(μ-NMe)₂P(μ₂-NMe)₂(μ₂-OMe₂)-[Li(OMe₂)]₂] } (**11b**) (Table 4).



Complexes **10** and **11a** have been characterized by multinuclear (¹H, ⁷Li, ¹³C, and ³¹P) NMR spectroscopy. The ¹H NMR spectrum of **10** in *d*₈-THF consists of two ^tBu

signals in a 1:2 ratio at 1.32 and 1.30 ppm, as well as singlets at 0.08 ppm and –0.94 ppm attributed to the SiMe₃ and AlMe₂ groups, respectively; the corresponding signals are observed in the ¹³C NMR spectrum. While the expected singlet is observed at 10.0 ppm in the ³¹P NMR spectrum, two resonances of approximately equal intensities are present in the ⁷Li NMR spectrum at 1.44 and –2.71 ppm. The latter observation suggests that, in THF solution, solvation of one lithium cation occurs to produce the solvent-separated ion pair { [Li(THF)₄]⁺{ Me₂Al[(μ-N^tBu)(μ-NSiMe₃)P(μ-N^tBu)₂]-Li(THF)₂ }⁻ (**12**), comprised of a spirocyclic monoanion and a tetrasolvated lithium counterion. The related tetra(naphthylimido)phosphate { [Li(THF)₄]⁺{ Li(THF)₂[(μ-Nnaph)₂P(μ-Nnaph)₂Li(THF)₂]⁻ is known to have a spirocyclic structure in the solid state.⁵

In contrast, the NMR data obtained for **11a** in *d*₈-THF are consistent with its solid-state structure. The inequivalence of the two types of *tert*-butyl groups is not apparent from the ¹H NMR spectrum, since only a single ^tBu resonance (1.34 ppm) is observed, accompanied by a singlet at –0.96 ppm attributed to the Me₂Al group. However, the ¹³C NMR spectrum clearly shows two sets of ^tBu signals (50.90 and 36.41 ppm; 50.78 and 34.45 ppm) with approximately equal intensities. The ⁷Li NMR spectrum exhibits a singlet at 1.23 ppm, indicating the equivalence of the two lithium cations and, thus, that formation of a solvent-separated ion pair does not occur for **11a**.

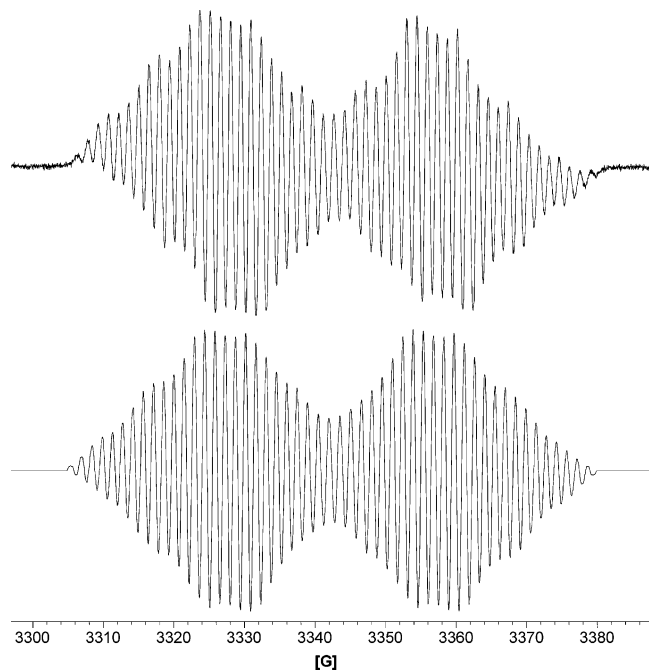
Reactions of **3a** with AlMe₃ and ⁿBuLi in THF are readily monitored by ³¹P NMR spectroscopy, as a downfield shift is observed with each subsequent deprotonation of the ligand. The signal at –15 ppm for (^tBuNH)₃PNSiMe₃ disappears as the dimethylaluminum complex **4** forms (5.6 ppm). Initially, the addition of ⁿBuLi results in the appearance of a peak at 8.2 ppm (**9a**); as more *n*-butyllithium is added, a signal for the dilithiated complex **12** is observed at 10.0 ppm. A similar trend in ³¹P NMR chemical shifts is observed when **3a** is reacted with 1, 2, and 3 equiv of *n*-butyllithium.¹¹

EPR and DFT Characterization of the Persistent Radicals { Me₂Al[(μ-NR)(μ-N^tBu)P(μ-N^tBu)₂Li(THF)₂][•] (**13a**, R = SiMe₃; **14a**, R = ^tBu). When exposed to air, THF solutions of **12** turn a bright blue color, whereas solutions of **11a** become a deep teal color indicative of the formation of radicals. In earlier studies, we have isolated stable cubic tetraimidophosphate radicals by oxidation of the dimeric lithium salt { Li₃[P(N^tBu)₃(NSiMe₃)]₂ } with 1 equiv of iodine or bromine.²⁵ Oxidations of the dimethyl-

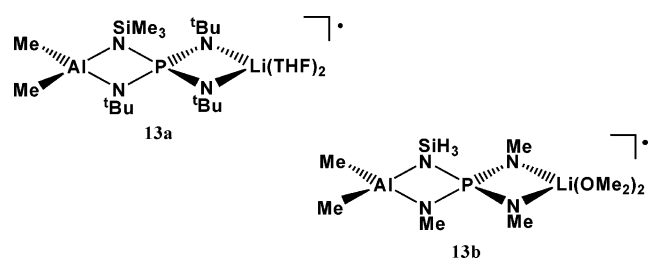
Table 4. Selected Bond Lengths (Å) and Angles (deg) for **11**

atoms	X-ray data for 11a	calcd for 11b	atoms	X-ray data for 11a	calcd for 11b
P(1)–N(1)	1.668(9)	1.661	Li(1)–O(1)	2.03(2)	1.960
P(1)–N(2)	1.671(9)	1.661	Li(1)–O(2)	2.15(2)	2.126
P(1)–N(3)	1.645(9)	1.668	Li(2)–O(2)	2.09(2)	2.126
P(1)–N(4)	1.597(10)	1.668	Li(2)–O(3)	2.00(2)	1.960
Al(1)–N(1)	1.881(10)	1.918	N(1)–Al(1)–N(2)	77.4(4)	76.7
Al(1)–N(2)	1.877(10)	1.918	N(3)–P(1)–N(4)	91.7(5)	93.2
Li(1)–N(3)	2.07(2)	2.009	N(1)–P(1)–N(2)	89.4(5)	91.5
Li(2)–N(3)	2.01(2)	2.009	Li(1)–N(3)–P(1)	91.0(7)	86.4
Li(1)–N(4)	2.13(2)	2.009	Li(2)–N(4)–P(1)	91.1(7)	86.4
Li(2)–N(4)	1.98(2)	2.009	Li(1)–O(2)–Li(2)	67.1(8)	66.4

aluminum complexes **11a** and **12** are expected to yield neutral spirocyclic radicals via the elimination of 1 equiv of THF-solvated lithium halide (Scheme 1). In practice, the addition of 0.5 equiv of iodine to a colorless solution of **12** in THF results in an intense blue color, which persists for several hours in solution. In the solid state, this mixed heterobimetallic tetraimidophosphate radical **13a** is stable for only a few minutes; however, it has been identified by its solution EPR spectrum (Figure 3). The observed pattern is

**Figure 3.** Experimental (top) and simulated (bottom) EPR spectra of **13a** in THF.

primarily a doublet due to coupling of the unpaired electron to the phosphorus atom (^{31}P , 100%, $I = 1/2$), with a total of approximately 50 lines due to additional hyperfine coupling to the nitrogen (^{14}N , 99%, $I = 1$), lithium (^7Li , 92.5%, $I = 3/2$), and aluminum (^{27}Al , 100%, $I = 5/2$) nuclei.



To aid in the interpretation of this complex spectrum, the model system $\{\text{Me}_2\text{Al}[(\mu\text{-NMe})(\mu\text{-NSiH}_3)\text{P}(\mu\text{-NMe})_2]\text{Li}(\text{Me}_2\text{O})_2\}^\bullet$ (**13b**) was examined by means of DFT calculations. The optimized molecular structure displays C_s symmetry, with the AlN_2P and LiN_2P rings perpendicular to one another. The two NMe groups that chelate the lithium cation have equivalent P–N bond lengths of 1.668 Å, while slightly shorter P–N bond distances are observed for the remaining P–NMe unit (1.655 Å) and the P–NSiH₃ (1.650 Å) fragment. The four metal–nitrogen distances are quite similar, with an average N–Al bond length of 1.958 Å and identical Li–N distances (2.030 Å).

Single-point calculations on **13b** indicated that the singly occupied molecular orbital (SOMO) is primarily a linear combination of nitrogen p orbitals (Figure 4). For symmetry reasons, the contributions from the nitrogen atoms in the AlN_2P and LiN_2P rings are not equivalent; a Mulliken population analysis of the atomic spin densities revealed that the spin density is concentrated on the nitrogen atoms of the LiN_2P ring. Although the SOMO has only minor contributions from phosphorus, lithium, and aluminum, these atoms have nonzero spin densities owing to spin polarization effects. In consequence, the EPR spectrum of **13b**, and hence that of its parent system **13a**, is expected to show large hyperfine couplings to one phosphorus and two equivalent nitrogen atoms. Smaller couplings to the other two nitrogen atoms and the metal centers should also be present and are responsible for the extensive hyperfine structure seen in the experimental EPR spectrum. While the calculations indicate that the Me_2Al^+ cation also contributes to the SOMO, coupling to the methyl protons is expected to be so small as to be unobservable.

An excellent simulation of the experimental EPR spectrum was obtained by including hyperfine couplings to one phosphorus atom (29.29 G), one lithium cation (1.43 G), one aluminum atom (1.39 G), a pair of equivalent nitrogen nuclei (7.20 G), and two unique nitrogen atoms (1.61 and 0.27 G) (Figure 3). These hyperfine couplings are consistent with the retention of the spirocyclic arrangement of **12** upon oxidation and the formation of the neutral radical $\{\text{Me}_2\text{Al}[(\mu\text{-N}^t\text{Bu})_2\text{P}(\mu\text{-N}^t\text{Bu})_2]\text{Li}(\text{THF})_2\}^\bullet$ (**13a**) (Scheme 1). To confirm the relative magnitudes of the HFCCs used to create the spectral simulation, the calculated data for **13b** were examined more closely. The differing spin densities on the

(25) Armstrong, A.; Chivers, T.; Parvez, M.; Boeré, R. *Angew. Chem., Int. Ed.* **2004**, *43*, 502–505.

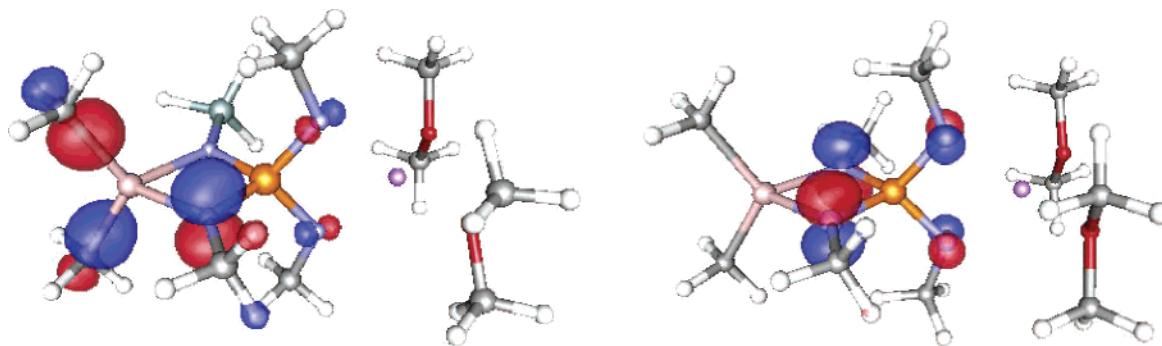


Figure 4. SOMO of **13b** (left) and **14b** (right).

Table 5. Experimental and Calculated HFCCs (G) for **13** and **14**

nucleus	spin	13a ^a	13b ^b	14a ^a	14b ^b
²⁷ Al	5/2	1.39	-1.22	1.99	-2.03
⁷ Li	3/2	1.43	-1.67	1.40	-1.58
³¹ P	1/2	29.29	-30.72	28.45	-30.96
¹⁴ N ⁱ Bu...Li	1	7.20	6.61	6.44	5.76
¹⁴ N ⁱ Bu...Al	1	1.61	-1.29	1.67	0.50
¹⁴ NSiMe ₃	1	0.27	-0.07		

^a Values used to simulate the experimental spectrum. ^b Values calculated for the model systems.

nitrogen atoms are reflected in the ¹⁴N HFCCs, which were calculated to be largest for the two equivalent NⁱBu ligands that chelate the lithium cation (6.61 G), smallest for the NSiMe₃ group (-0.07 G), and of an intermediate size for the unique NⁱBu nitrogen atom (-1.29 G). The calculated HFCCs are shown in Table 5, along with the values used to create the spectral simulation shown in Figure 3. The largest calculated HFCC was to the phosphorus nucleus (-30.72 G), while the much smaller aluminum (-1.22 G) and lithium (-1.67 G) couplings are calculated to be fairly similar. These calculated HFCCs provide good support for the proposed spirocyclic structure of the radical **13a**.

The one-electron oxidation of **11a** with halogens yields the persistent teal radical {Me₂Al[(μ-NⁱBu)₂P(μ-NⁱBu)₂Li(THF)₂]}• (**14a**) with a stability similar to that of **13a**: it persists in solution for approximately 3 h but cannot be isolated in the solid state. The solution EPR spectrum of this radical is shown in Figure 5. DFT calculations were carried out on the model system {Me₂Al[(μ-NMe)₂P(μ-NMe)₂Li(Me₂O)₂]}• (**14b**). The geometry-optimized structure of this molecule has C_{2v} symmetry, with the two MN₂P (M = Li, Al) rings perpendicular to one another. The two pairs of P-N bond lengths in this structure are equivalent within the accuracy of the calculations, at 1.650 and 1.653 Å for the P-N-Al and P-N-Li distances, respectively. As was seen for **13b**, the Al-N bonds (1.936 Å) are slightly shorter than the Li-N bonds (1.984 Å).

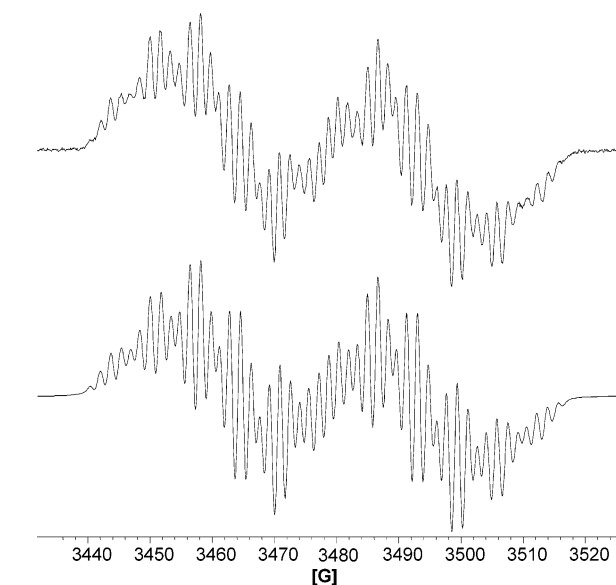
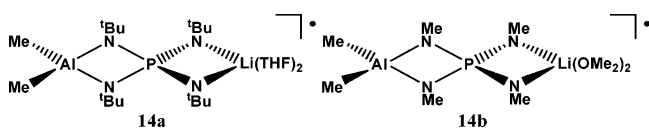


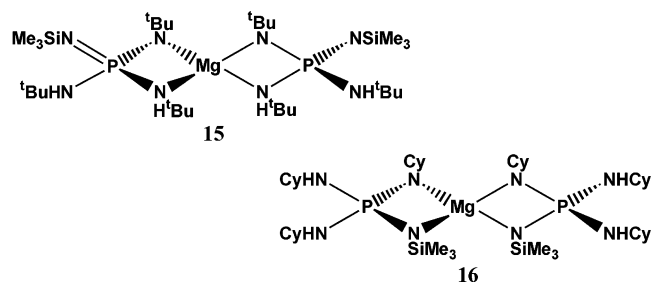
Figure 5. Experimental (top) and simulated (bottom) EPR spectra of **14a** in THF.

spectrum was obtained by including hyperfine couplings to two pairs of equivalent ¹⁴N atoms (6.44 G, 1.67 G), one lithium cation (1.40 G), one dimethylaluminum cation (1.99 G), and the phosphorus atom (28.45 G). The close agreement of the experimental and calculated HFCCs, which are shown in Table 5, confirms that the oxidation product of **11a** is indeed the neutral spirocycle radical {Me₂Al[(μ-NⁱBu)₂P(μ-NⁱBu)₂Li(THF)₂]}• (**14a**). Unlike other related tetraimidophosphate radicals,²⁵ the EPR spectra of **13a** and **14a** are not dependent on sample concentration, indicating that dissociation of these spirocyclic radicals (e.g., by solvation of the lithium cations) does not occur even in very dilute solutions.

Synthesis and Characterization of {Mg[(μ-NⁱBu)-(μ-NHⁱBu)P(NHⁱBu)(NSiMe₃)₂]}₂ (15**), {Mg[(μ-NCy)(μ-NSiMe₃)P(NHCy)₂]}₂ (**16**), {Mg[(μ-S)(μ-NⁱBu)P(NHⁱBu)₂]}₂ (**17**), and {[MgOH][(μ-S)(μ-NⁱBu)P(NHⁱBu)₂]}₆ (**18**).** The reaction of MgBu₂ with 2 equiv of (t-BuNH)₃PNSiMe₃ (**3a**) produces a complex in which a magnesium dication is bis-chelated by two imidophosphate monoanions {Mg[(μ-NⁱBu)-(μ-NHⁱBu)P(NHⁱBu)(NSiMe₃)₂]}₂ **15**. This compound has been characterized by multinuclear NMR spectroscopy and displays a single resonance in its ³¹P spectrum at 5.2 ppm. The ¹H NMR spectrum of **15** in C₆D₆ is interesting in that it consists of three t-Bu signals of equal intensities, one SiMe₃

Further examination of the data indicated that the composition of the SOMO (Figure 4) of **14b** is similar to that of **13b**. An excellent simulation of the experimental EPR

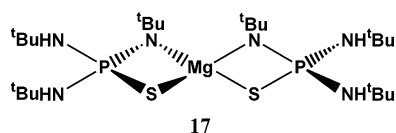
resonance, and two broad doublets that are attributable to two inequivalent NH protons. These data indicate that the imidophosphate ligand is coordinated to the magnesium atom through one NH^tBu and one N^tBu group. This is somewhat surprising, considering the well-established structures of the dimethylaluminum complexes **4**, **9a**, **10**, and **11a**, in which the aluminum atom is chelated by the two imido groups, while the amido substituents remain noncoordinated. This preference is attributed to the fact that Mg²⁺ is a harder acid than Me₂Al⁺ and will, therefore, prefer coordination to the harder base NH^tBu over coordination to the NSiMe₃ nitrogen atom.



The reaction of MgBu₂ with the cyclohexyl derivative (CyNH)₃PNSiMe₃ (**3b**) in a 1:2 molar ratio produces a complex that was identified as {Mg[(μ-NCy)(μ-NSiMe₃)P(NHCy)₂]₂} (**16**) on the basis of NMR data and CHN analysis. The appearance of resonances for only two cyclohexyl environments in a 2:1 ratio in the ¹³C NMR spectrum (in C₆D₆) suggests coordination of two imido groups (NSiMe₃ and NCy) to the magnesium center and two exocyclic cyclohexylamido groups.

Attempts to effect further deprotonation by changing the reaction conditions and stoichiometry were unsuccessful; the only products observed by ³¹P NMR were the monoanionic complexes **15** and **16**. Similar low reactivities were observed for these species toward dimethylzinc and trimethylaluminum. However, the tris(amino)(thio)phosphate SP(NH^tBu)₃ (**2a**) was previously found to be slightly more reactive than (RNH)₃PNSiMe₃ toward ZnMe₂.⁷ Consequently, reactions between MgBu₂ and **2a** were explored using various stoichiometries and reaction conditions.

When 0.5 equiv of MgBu₂ was employed, a white powder was recovered that gave rise to only one ³¹P NMR signal (38.8 ppm). The ¹H NMR and ¹³C NMR spectra of the product in C₆D₆ indicated the presence of two inequivalent ^tBu environments. The 1:2 integration of the ¹H NMR peaks (at 1.56 and 1.45 ppm), and the absence of any resonances attributable to a BuMg⁺ cation, is consistent with the formation of the *N,S*-chelated complex {Mg[(μ-S)(μ-N^tBu)P(NH^tBu)₂]₂} (**17**).



Attempts to grow single crystals of **17** from a variety of organic solvents were unsuccessful; however, it was found

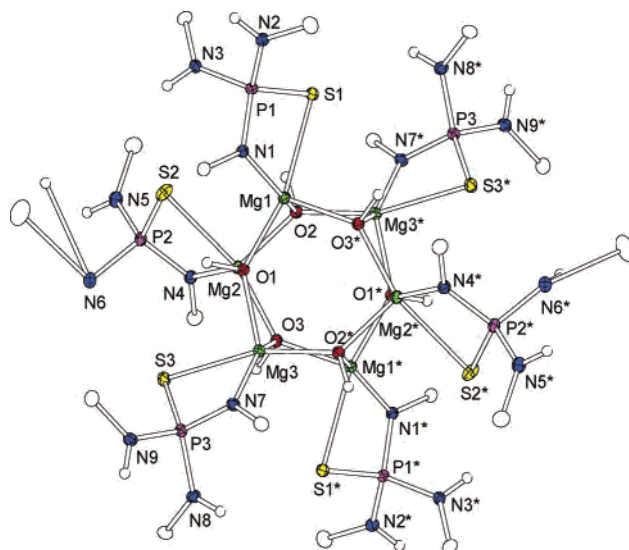


Figure 6. Thermal ellipsoid diagram of **18** (30% probability ellipsoids). Only α -carbons of the *tert*-butyl groups are shown.

Table 6. Selected Bond Lengths (Å) and Angles (deg) for **18**^a

P(1)–N(1)	1.6078(16)	P(1)–N(2)	1.6532(17)
P(1)–N(3)	1.6721(18)	P(1)–S(1)	2.0112(8)
Mg(1)–O(1)	2.0266(15)	Mg(1)–O(2)	2.1191(15)
Mg(1)–O(3)	2.0030(16)	Mg(2)–O(1)	2.1198(15)
Mg(2)–O(2)	2.0042(15)	Mg(2)–O(3)	2.0325(15)
Mg(3)–O(1)	2.0039(15)	Mg(3)–O(2)	2.0346(16)
Mg(3)–O(3)	2.1208(15)	Mg(1)–N(1)	2.0675(17)
Mg(1)–S(1)	2.6290(10)	Mg(2)–N(4)	2.0686(18)
Mg(2)–S(2)	2.6166(11)	Mg(3)–N(7)	2.0677(18)
Mg(3)–S(3)	2.6237(11)	Mg(1)–Mg(2)	3.0317(12)
N(2)–P(1)–N(3)	100.60(9)	P(1)–S(1)–Mg(1)	75.56(3)
P(1)–N(1)–Mg(1)	102.69(8)	S(1)–Mg(1)–N(1)	74.47(5)
O(1)–Mg(1)–O(2)	83.08(6)	Mg(1)–O(1)–Mg(2)	93.94(6)
S(1)–P(1)–N(1)	104.74(6)	O(1)–Mg(3)–O(2*)	103.21(6)
O(2)–Mg(2)–O(3)	102.40(6)	O(1)–Mg(2)–O(2)	83.60(6)

^a Symmetry transformations used to generate equivalent atoms: #1 $-x, -y + 1, -z$.

that prolonged storage of this complex in Et₂O leads to the of the hexameric complex {[Mgμ₃-OH][(μ-S)(μ-N^tBu)P(NH^tBu)₂]₆} (**18**), presumably via the partial hydrolysis of **17**. An X-ray structure of **18** (Figure 6) revealed that the core of this complex is composed of six [Mgμ₃-OH]⁺ cations linked together via Mg–O bonds to form two stacked six-membered Mg₃O₃ rings (or a Mg₆O₆ hexagonal prism). Each magnesium center is also *N,S*-chelated by the monoanionic ligands [SP(NH^tBu)₂(N^tBu)][−], rendering the Mg²⁺ cations pentacoordinate. The bis(amido)imido(thio)phosphate ligands radiate outward from the Mg₆O₆ core, with the two ^tBuNH groups of each ligand remaining noncoordinated. Selected bond lengths and bond angles for **18** are summarized in Table 6. The three Mg–O bonds in the center of the molecule are inequivalent: each magnesium atom has one short bond [e.g., Mg(1)–O(3) = 2.0030(16) Å], which can be considered a discrete [MgOH]⁺ unit; one slightly weaker bond [Mg(1)–O(1) = 2.0266(15) Å] to a second oxygen atom within a Mg₃O₃ ring; and a third, longer distance [Mg(1)–O(2) = 2.1191(15) Å] between the magnesium of one Mg₃O₃ ring and an oxygen atom in the other six-membered ring. The average Mg–O bond distance in **18** is 2.054(15) Å, compare 1.9473(13)–2.0020(14) Å for the related complex {[NaMg-

Table 7. Reactivities of EP(NHR)₃ toward Metal Alkyl Reagents MR_x

(E, R)	ⁿ BuLi	ZnMe ₂	AlMe ₃	MgBu ₂
O, ^t Bu	1- or 2-fold deprotonation ^a	3-fold deprotonation; multiple products formed ^b	1- or 2-fold deprotonation ^c	Possibly 3-fold deprotonation; multiple products formed ^d
S, ^t Bu	1- or 2-fold deprotonation; sulfur extrusion and dimerization to a metaphosphate ^e	monodeprotonation ^b	1- or 2-fold deprotonation ^d	monodeprotonation; sulfur extrusion at high temperatures ^d
NSiMe ₃ , Cy	3-fold deprotonation ^f	monodeprotonation ^b	dimerization to metaphosphate ^d	monodeprotonation; NCy, NSiMe ₃ chelation ^d
NSiMe ₃ , ^t Bu	3-fold deprotonation ^f	monodeprotonation; N ^t Bu, NSiMe ₃ chelation ^b	monodeprotonation; N ^t Bu, NSiMe ₃ chelation ^d	monodeprotonation; N ^t Bu, NH ^t Bu chelation ^d

^a References 6 and 8. ^b Reference 7. ^c Reference 12. ^d This work. ^e References 8 and 10. ^f Reference 6.

(N^tPr)₂(O)·THF₆, which also contains a core of two stacked Mg₃O₃ rings.²⁶ The bis(amido)imido(thio)phosphate ligand, as expected, contains one short P–N bond to the *tert*-butylimido nitrogen atom [P(1)–N(1) = 1.6078(16) Å], while longer distances are observed to the two *tert*-butyl-amido groups [P(1)–N(2) = 1.6532(17) Å; P(1)–N(3) = 1.6721(18) Å].

The deliberate hydrolysis of **17** with 1 equiv of H₂O was attempted in both diethyl ether and toluene. Reaction times of more than 10 min result in the complete hydrolysis of **17**, and a mixture of SP(NH^tBu)₃ (**2a**) and magnesium oxide was recovered. Over shorter reaction times, two signals of equal intensity attributable to **2a** and **17** are observed in the ³¹P NMR spectrum of the product mixture. It is thought that removal of the solvent under a vacuum facilitates the loss of water from **18** so that the hydroxide complex is not observed.

The reaction of **2a** and MgBu₂ in a 1:1 stoichiometry gave a complex mixture of products, as revealed by ³¹P NMR spectra. The marked increase in the number of products formed from the reaction of EP(NH^tBu)₃ (E = O, S) with dibutylmagnesium as compared to dimethylzinc⁷ can be attributed to the higher electrophilicity of the magnesium cations and their tendency to adopt octahedral rather than tetrahedral coordination modes.

(26) Drummond, A. M.; Gibson, L. T.; Kennedy, A. R.; Mulvey, R. E.; O'Hara, C. T.; Rowlings, R. B.; Weightman, T. *Angew. Chem., Int. Ed.* **2002**, *41*, 2382–2384.

Conclusions

The trisamidophosphates EP(NHR)₃ (E = O, S, NR'; R = alkyl) display similar reactivities toward AlMe₃ as they do toward ZnMe₂, with the highest reactivity observed for E = O and the lowest reactivity for E = NR'. Three-fold deprotonation of these species to form imido analogues of the well-known aluminophosphates does not occur. In the case of E = NR', two different reaction pathways can occur to yield imido analogues of *ortho*- or *meta*-phosphate, depending on the identity of the R group employed. Reactions of AlMe₃ with SP(NH^tBu)₃ yield *N,S*-chelated complexes. Lithiation of aluminum bis(amido)bis(imido)-phosphates produces heterobimetallic complexes that form persistent radicals upon oxidation with iodine. Reactions with MgBu₂ lead to multiple products for E = O and S, but only a simple monoanionic complex of E = NR'. The different behaviors of **1–3** toward various organometallic (metal alkyl) reagents are summarized in Table 7.

Acknowledgment. We thank the University of Calgary and the Natural Sciences and Engineering Research Council of Canada for funding and Helsingin Sanomain 100-vuotissäätiö for a scholarship (H.M.T.).

Supporting Information Available: Crystallographic CIF files are included in the Supporting Information. This material is available free of charge via the Internet at <http://pubs.acs.org>.

IC050689E